

ROTD 2/20/06

1) Types of strain: a) Torsional due to electrons in the bonds of adjacent atoms repelling each other (depends upon dihedral angle); b) Steric due to two things trying to share the same physical space; c) Angle due to deviation from ideal bond angle (violating VSEPR theory).

2) Butane is the smallest molecule that we can study for both steric and torsional strain. We can use what we know about butane conformations to study larger molecules (like cyclohexane).

3) Attractive forces make molecules stick together, and the strength of the interactions help determine some physical properties (melting point, boiling point, etc.).