

# **ROTD 2/9/06**

**1) Polar Covalent Bonds happen when atoms of differing electronegativity bond to each other.**

**2) Dipole Moment is a measure of how polar the bond is (i.e. how disproportionate the electron sharing is).**

**3) Lewis Structures are a way of combining Lewis Dot structures to represent molecules.**

**4) Polar molecules are molecules with polar bonds that give the whole molecule regions of higher and lower electron density.**